

①



\boxed{B}

②



in soli HCO e CO^-

$$m_{HCO} = 0.20$$

$$m_{NaOH} = 0.15$$

primo dopo la reazione:

$$m_{HCO} = 0.05 \text{ mol} \quad m_{CO^-} = 0.15 \text{ mol}$$

$$[HCO] = 0.05 M \quad [CO^-] = 0.15 M$$

1 LITRO

$$pH = pK_a + \log \frac{[CO^-]}{[HCO]}$$

$$= 7.96 + \log \frac{0.15}{0.05} = 7.96 + 0.47$$

$$= 8.44$$

\boxed{C}

1

3



32 mol of B_2H_6

B

4



$$\text{PM} = 193.88$$

$$\% \text{Fe} = 28.8 \% \quad \% \text{N} = 3 \frac{(14.01)}{\text{PM}} \cdot 100 = 21.7$$

$$\% \text{O} = 49.5 \%$$

A

5



$$n_{e^-} = \frac{Q}{F} = \frac{i \cdot t}{F} = \frac{1000 \cdot 12 \cdot 60}{96500}$$

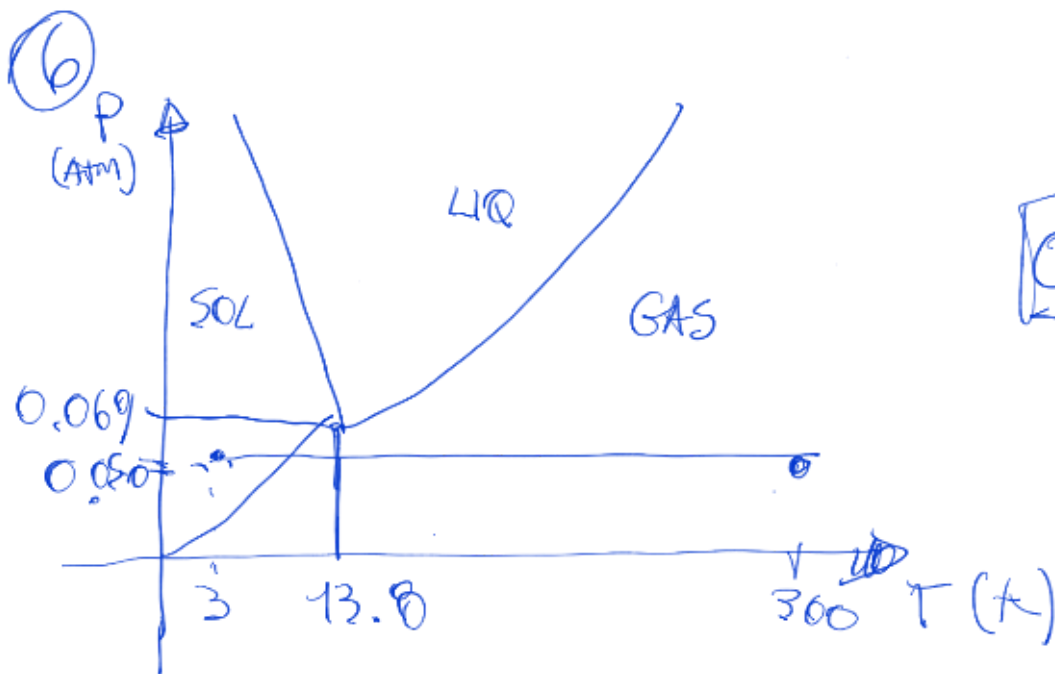
$$\approx 7.66 \text{ moli}$$

$$n_{Cu} = n_{e^-} \cdot \frac{1}{2} = 3.73 \text{ moli}$$

$$\text{MAGSA Cu} = 3.73 \cdot 63.55$$

$$\approx 237.1 \text{ g}$$

C



C