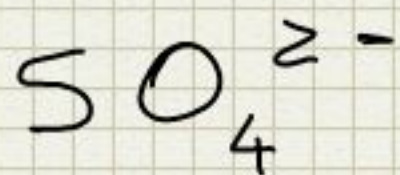
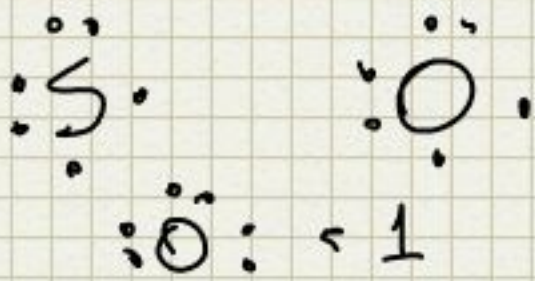


SOLFATO

(ACIDO SOLFORICO)

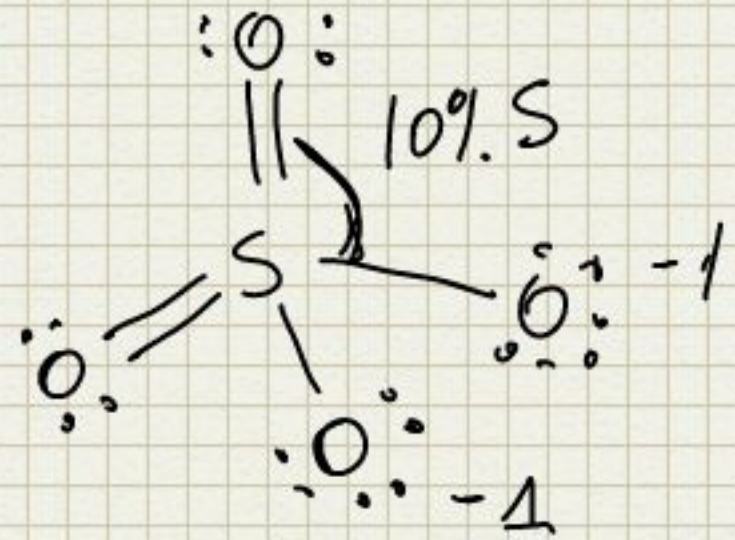
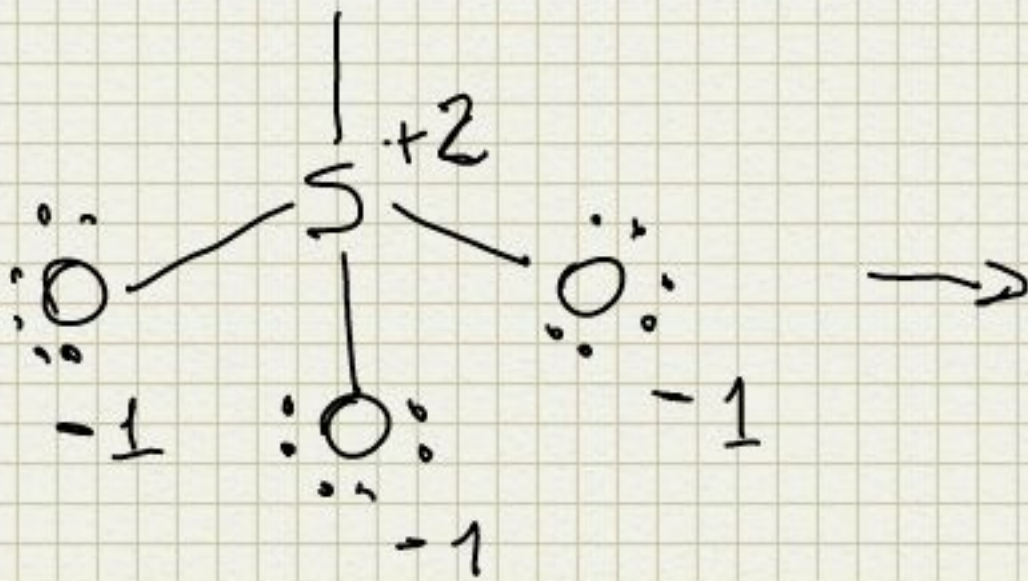


$$\text{m.o. S} = +6$$

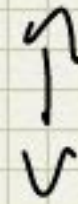


$$6 \cdot 5 + 2 = 32$$

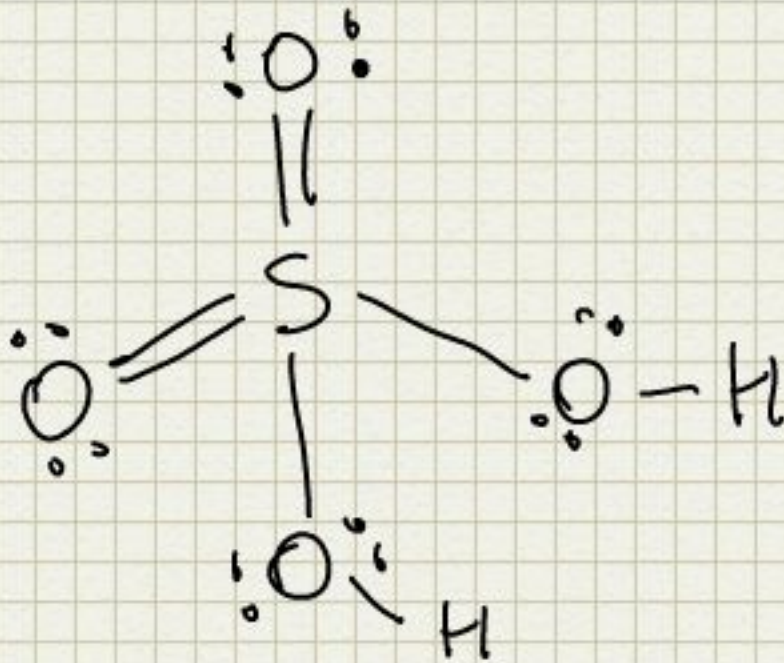
16 COPPIE



S IBRIDATO sp^3

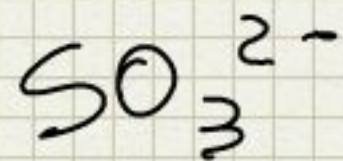


ACIDO SOLFORICO

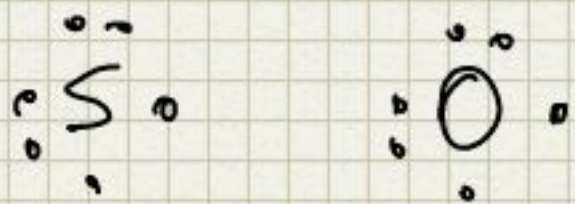


SOLFITO

(ACIDO SOLIFOROSO)

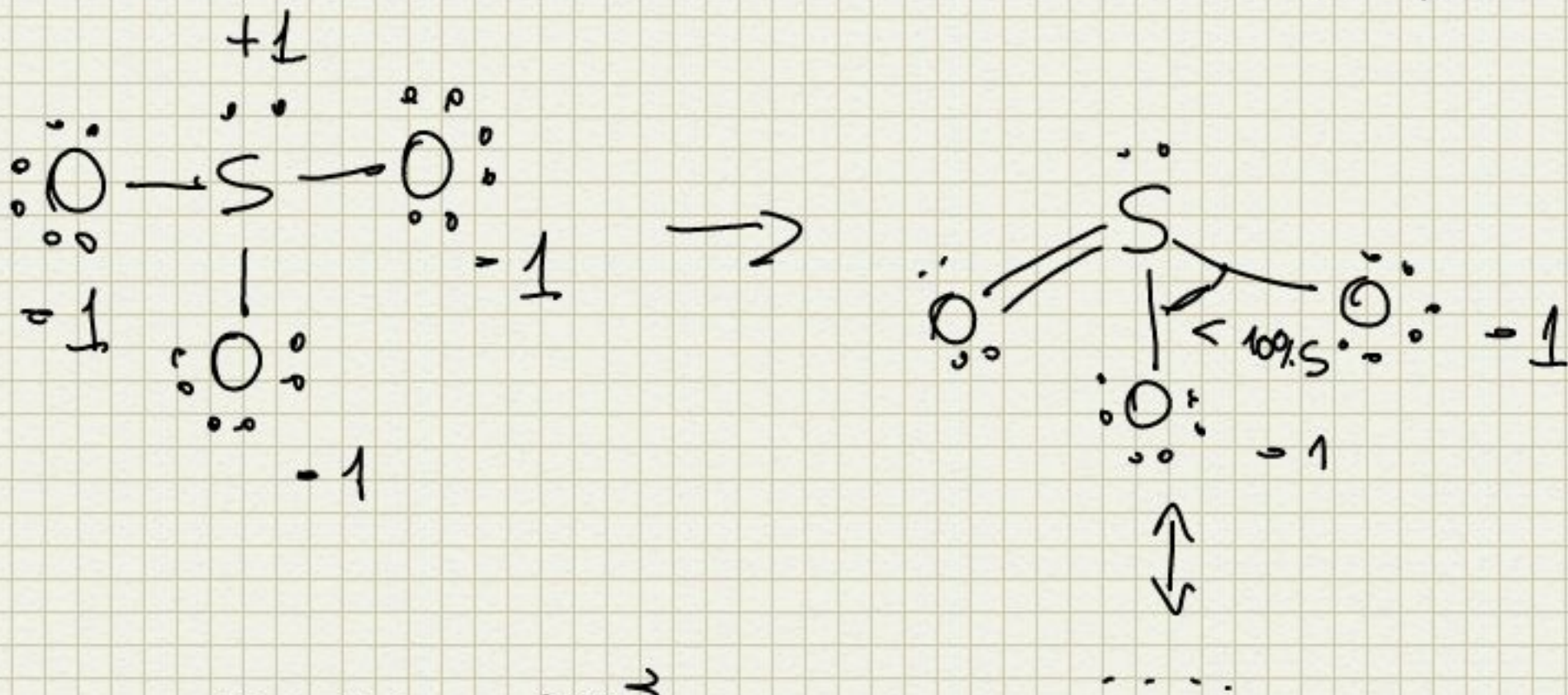


$$\text{N.OX. S} = +4$$



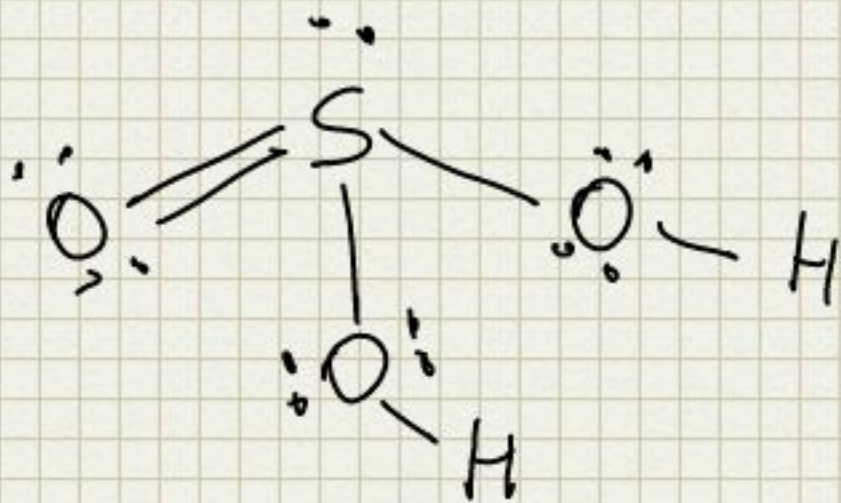
$$6 \cdot 4 + 2 = 26$$

13 COPPIE

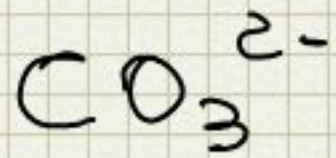


S IBRIDATO sp^3

ACIDO SOLFOROSO



CARBONATO (ACIDO CARBONICO)

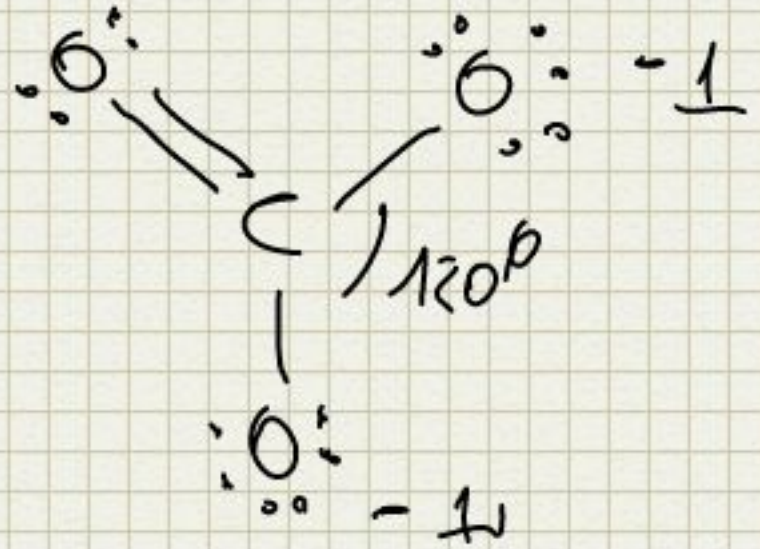
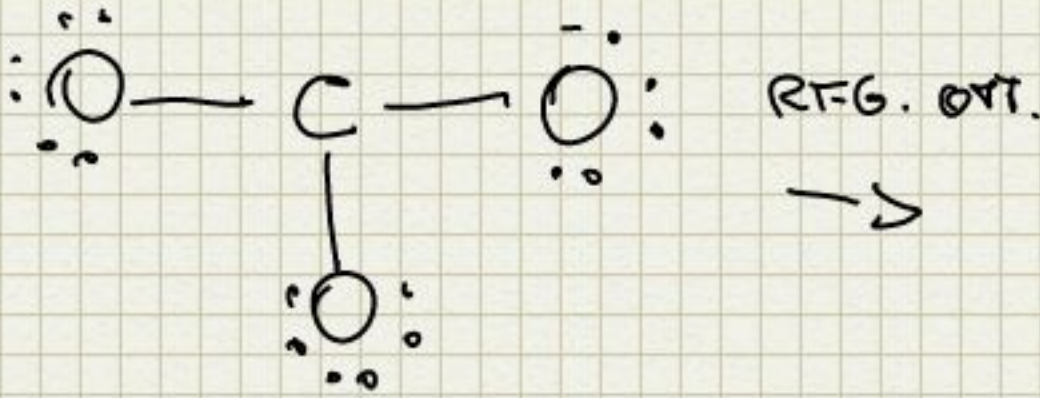


$$\text{n. OX C} = +4$$



$$6 \cdot 3 + 4 + 2 = 24$$

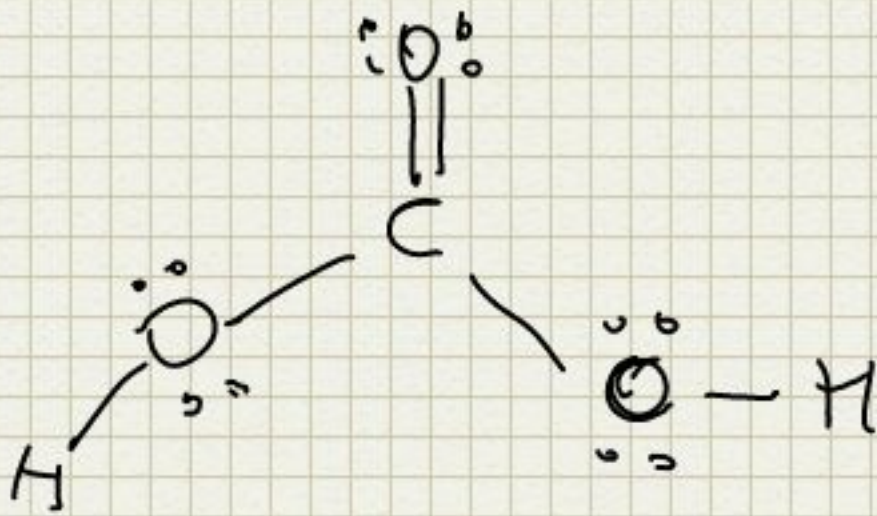
12 COPPIE



PLANARE

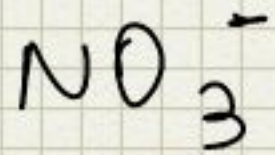
C IBRIDATO Δp^2

ACIDO CARBONICO

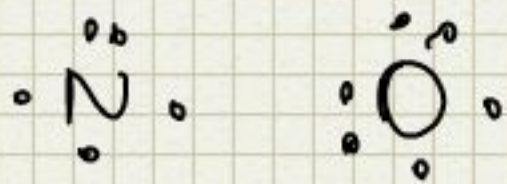


NITRATO

(ACIDO NITRICO)

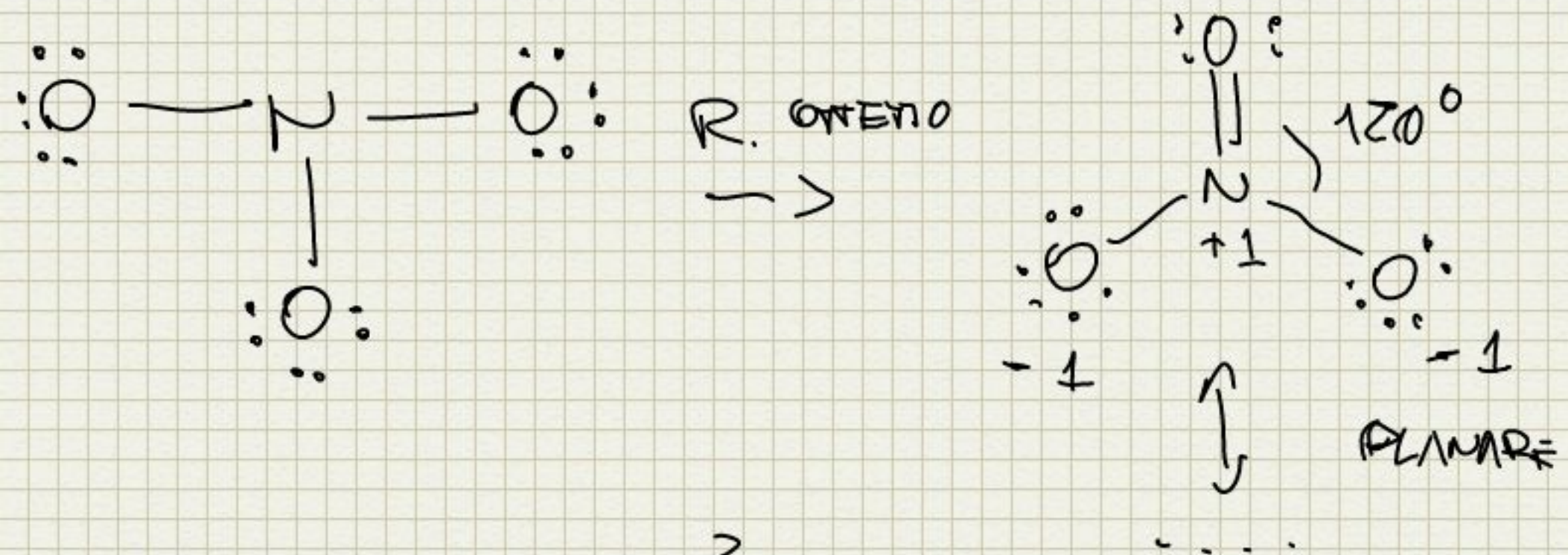


NOX, N = +5



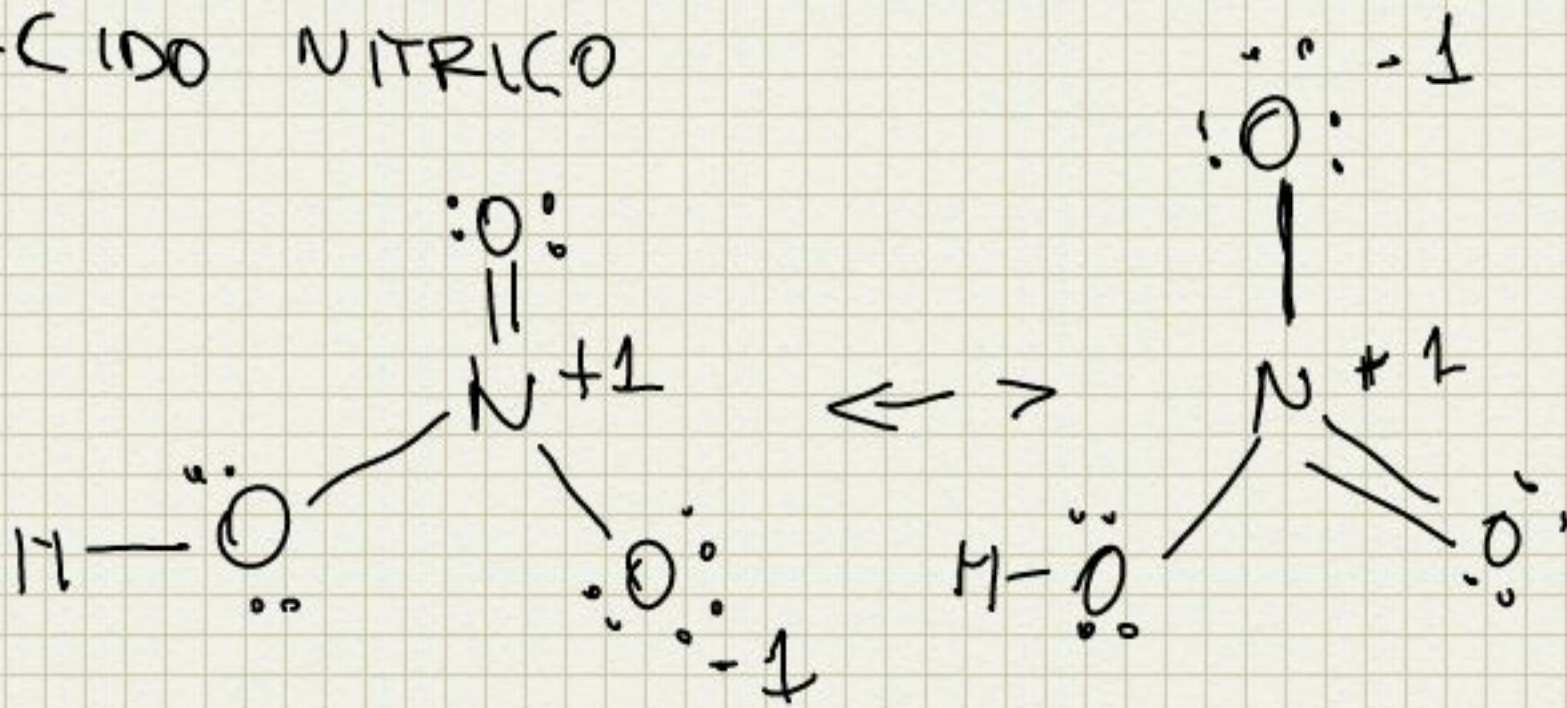
$$6 \cdot 3 + 5 + 1 = 24$$

12 COPPIE



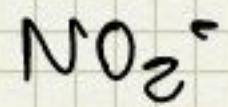
N IBRIDATO sp^2

ACIDO NITRICO

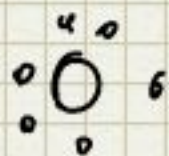
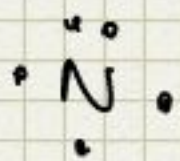


NITRITO

(ACIDO NITROSO)

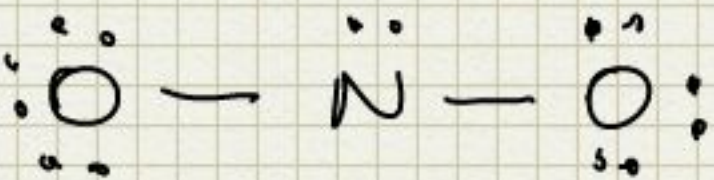


$$\text{n. ox N} = +3$$

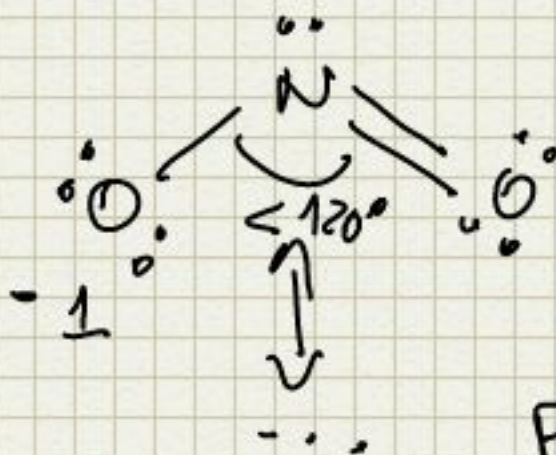


$$6 \cdot 2 + 5 + 1 = 18$$

di COPPIE

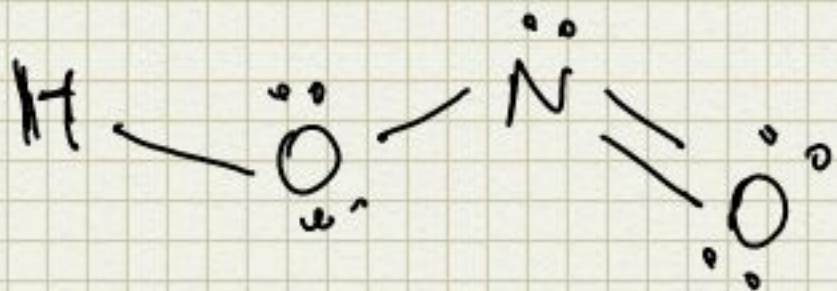


R. OTT.
→



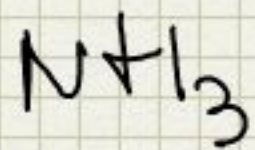
N IBRIDATO sp^2

ACIDO NITROSO



AMMONIACA

(IONE AMMONIO)

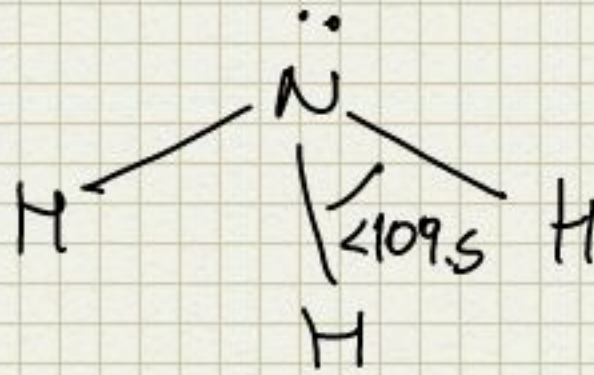
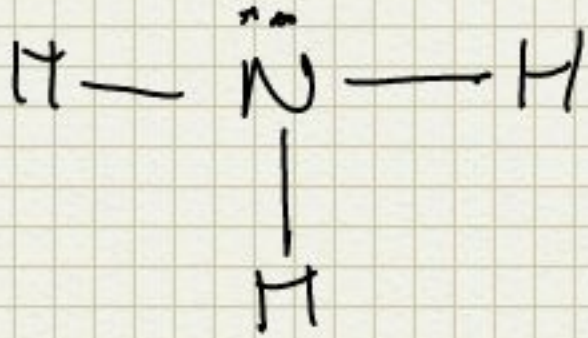


N. ox. N = -3



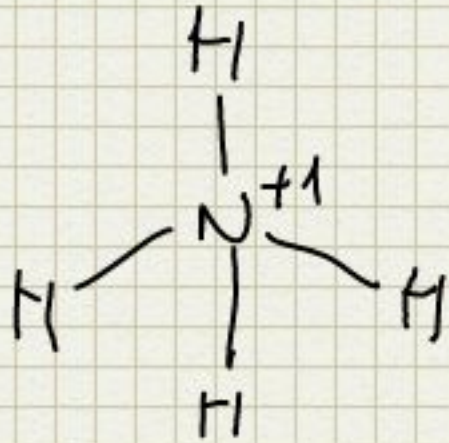
$$5 + 3 = 8$$

4 COPPIE

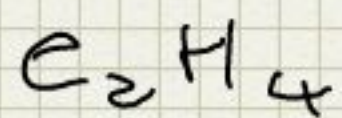


N IBRIDATO sp^3

IONE AMMONIO



ETHYLENE



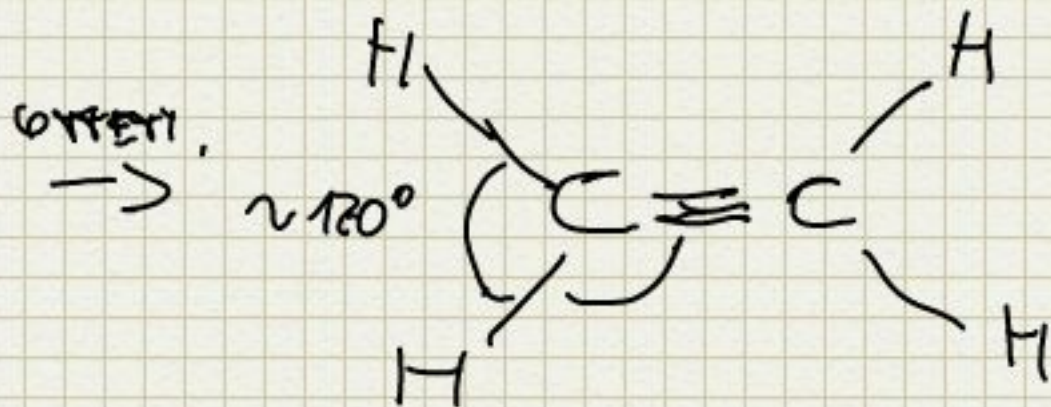
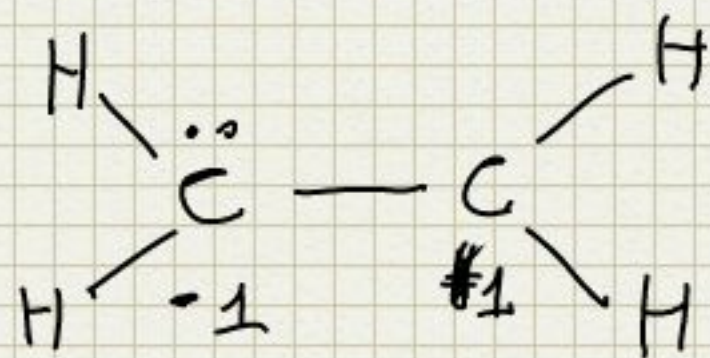
M.OX

C = -2



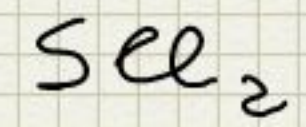
$4 \cdot 2 + 4 = 12$

6 COPPIE

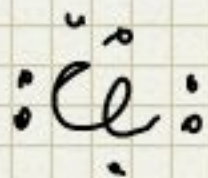
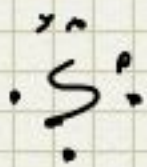


C IBRIDAZIONE sp^2 e legamo π sovrapp. $2p$

DICLORURO DI ZOLFO

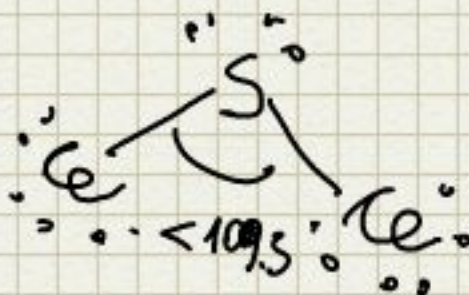
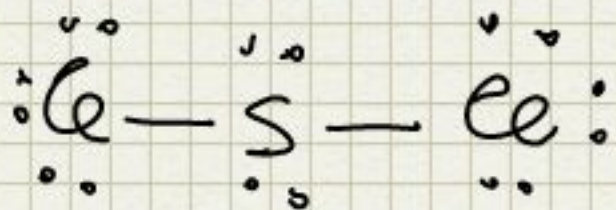


m. ox S, + 2



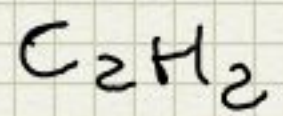
$$7 \cdot 2 + 6 = 20$$

10 COPPIE



S IBRIDATO sp^3

ACETYLENE

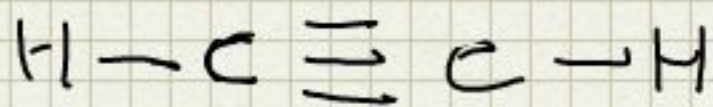


$$MOX. \quad C = -1$$



$$4 \cdot 2 + 2 = 10$$

5 COPPIE

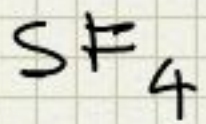


C IBRIDATO sp

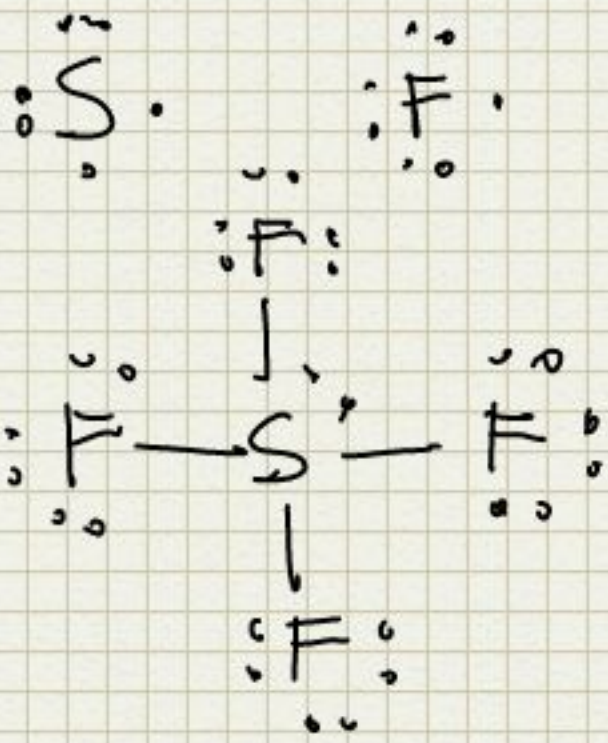
1 σ Δ ovrof. sp

2 π Δ ovof. $2p$

TETRAFLUORURO DI ZOLFO



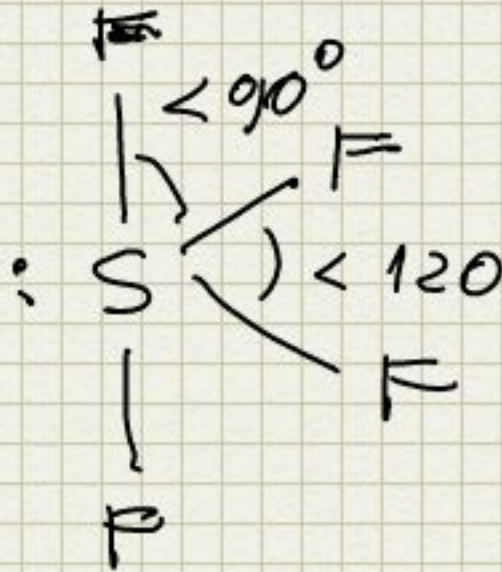
m.ox. S = +4



$$7 \cdot 4 + 6 = 34$$

17 COPPIE

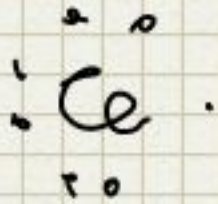
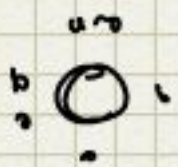
S IBRIDATO sp^3d



IPOLORTO

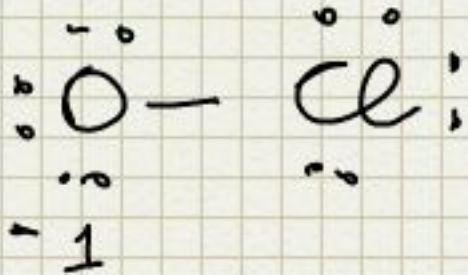
CeO^+ m.ox $Ce = +1$

ACIDO IPOLOROSO H_2CeO_4



$$7 + 6 + 1 = 14$$

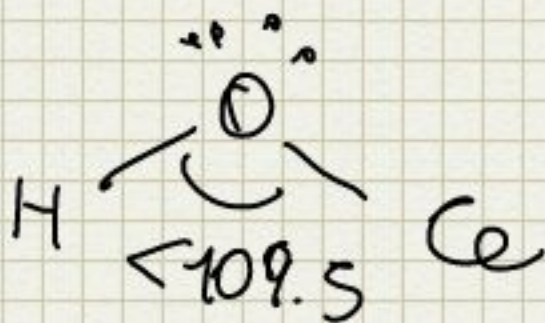
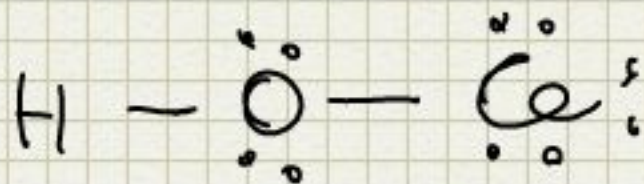
7 coppie



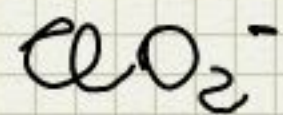
lineare

ACIDO

IPOCLOUROSO

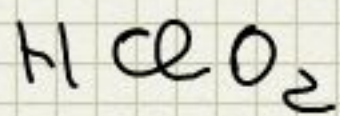


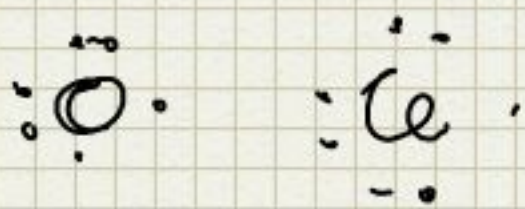
CLORITO



n. ox Cl = +3

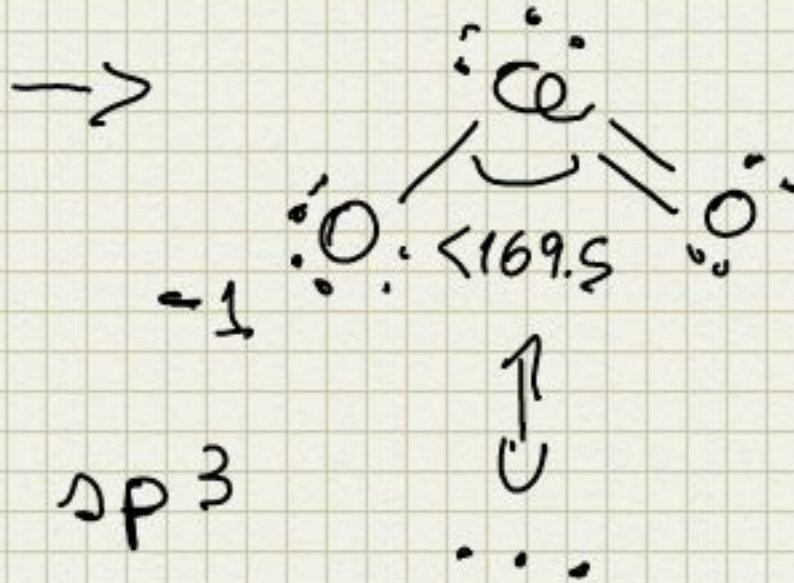
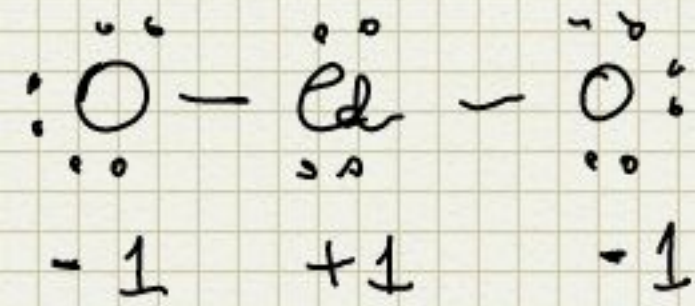
ACIDOO CLOROSO





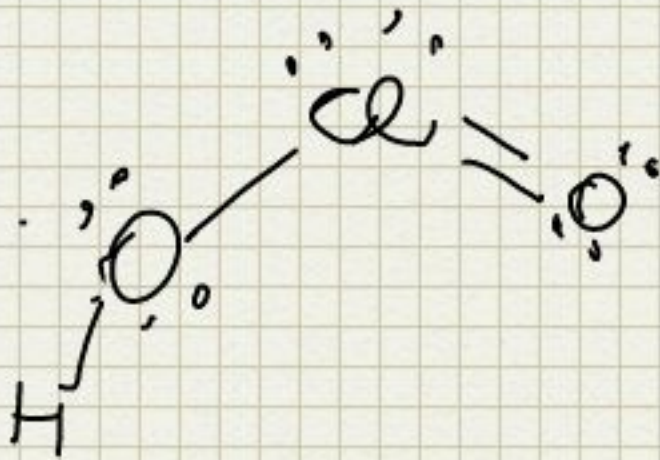
$$6 \cdot 2 + 7 + 1 = 20$$

10 COPPIE

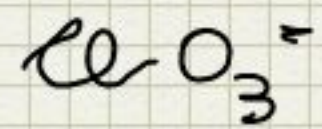


Cl IBRIDATO sp^3

ACIDO CLOROSO



CLORATO



num. ox Cl = +5

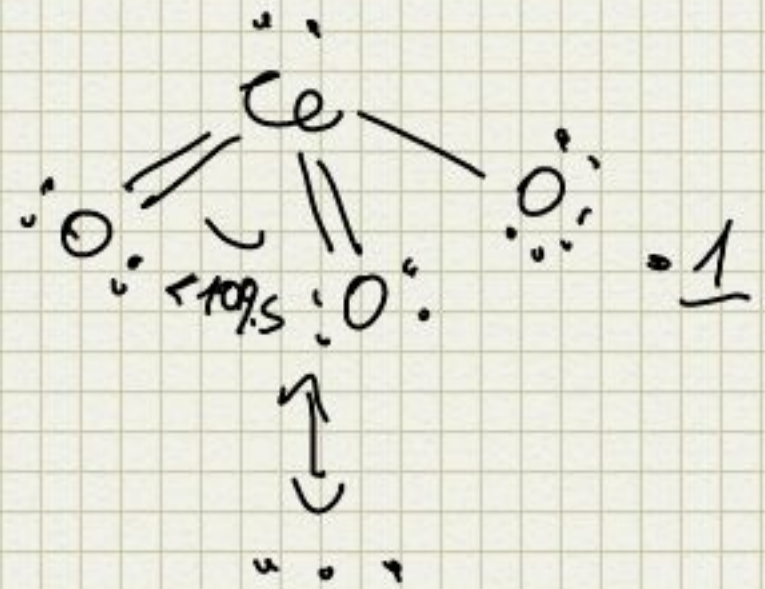
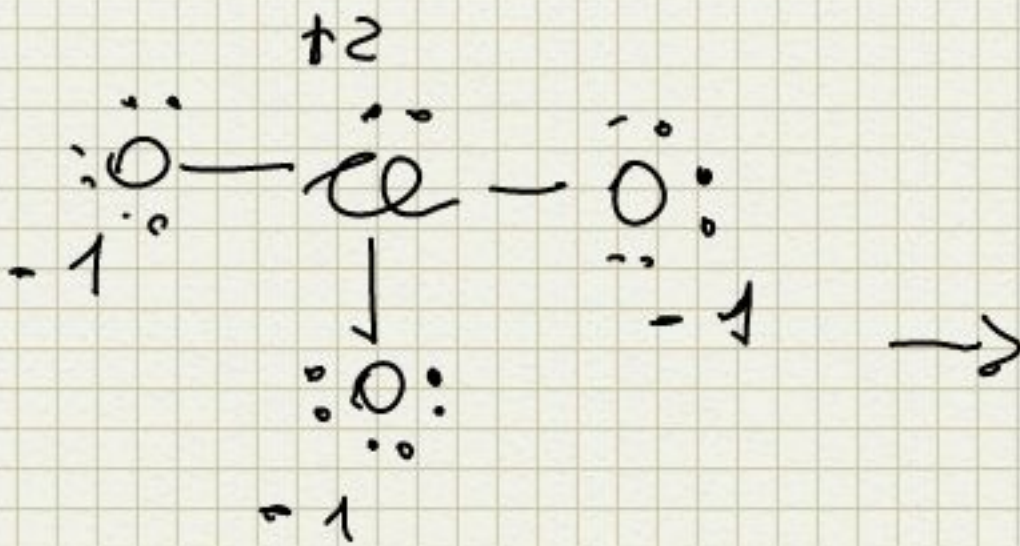
ACIDO CLORICO



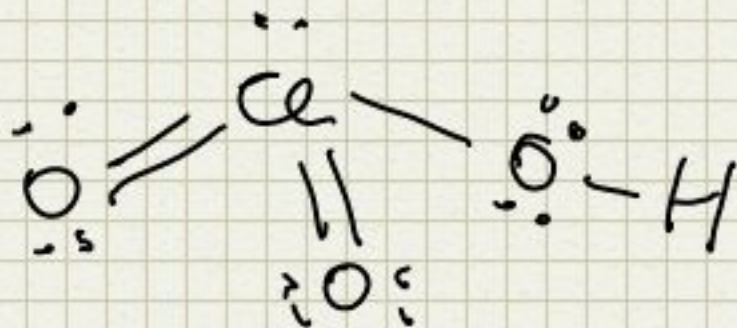


$$6 \cdot 3 + 7 + 1 = 26$$

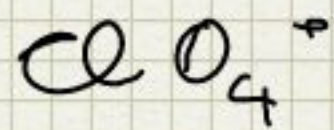
13 COPPIE



Cl IBRIDATO sp^3
ACIDO CLORICO

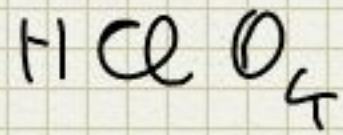


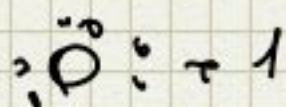
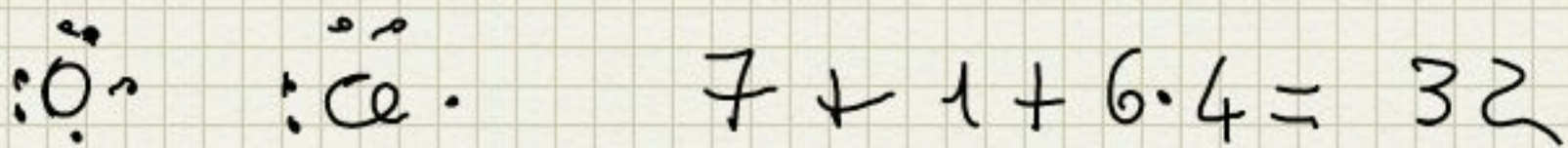
PERCLORATO



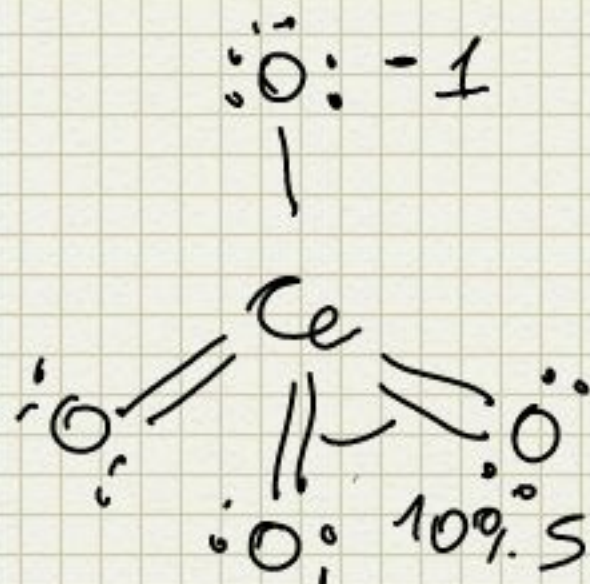
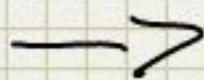
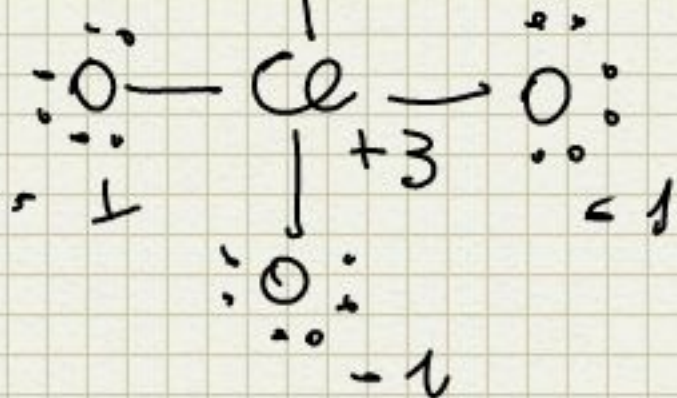
n. ox. Cl = +7

ACIDO PERCLORICO

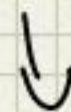




16 COPPIE

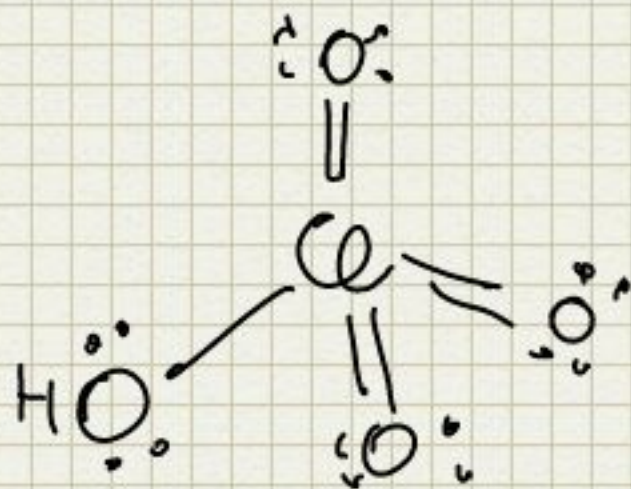


ce IBRIDATO sp^3



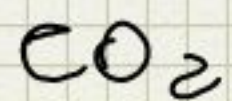
...

ACIDO PERCLORICO



ANIDRIDE

CARBONICA

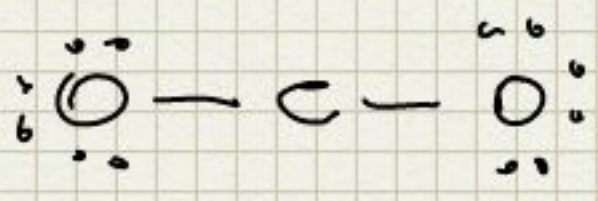


$$\text{max } C = +4$$

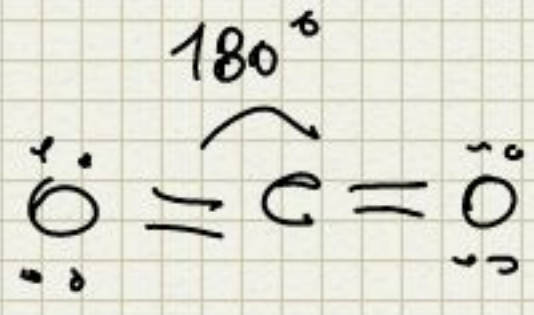


$6 \cdot 2 + 4 = 16$

8 COPPIE



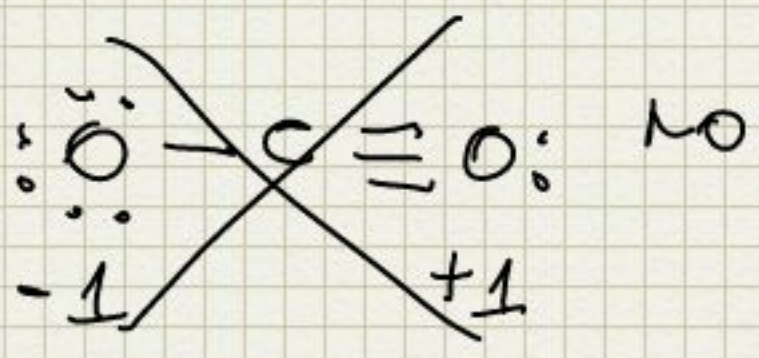
OTTE.
→



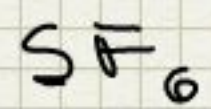
C IBRIDATO SP

2 legami σ

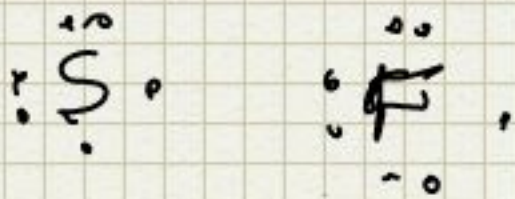
2 legami π



ESAFLUORURO DI ZOLFO

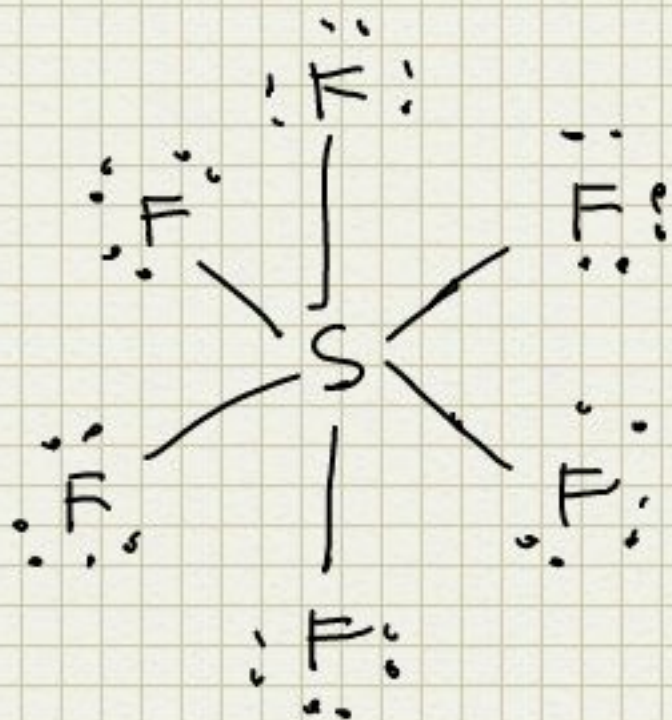


MOX $\xi = +6$



$$7 \cdot 6 + 6 = 48$$

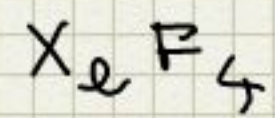
24 COPPIE

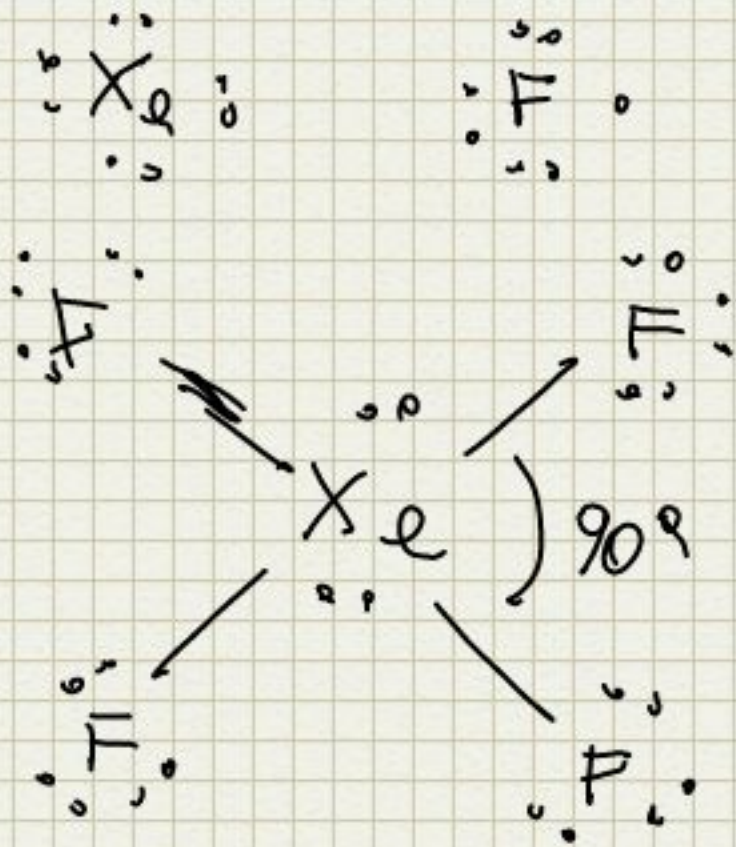


§ IBRIDATO sp^3d^2

$$\hat{R}SF = 90^\circ$$

TETRAFLUORURO DI XENO





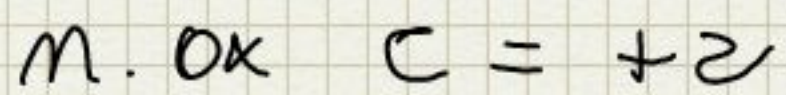
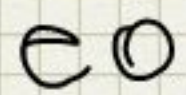
$$7 \cdot 4 + 8 = 36$$

18 COPPIE

Xe IBRIDATO sp^3d^2

PLANARE

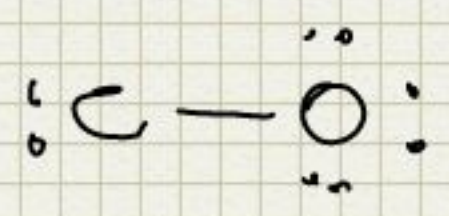
MONOSSIDO DI CARBONIO



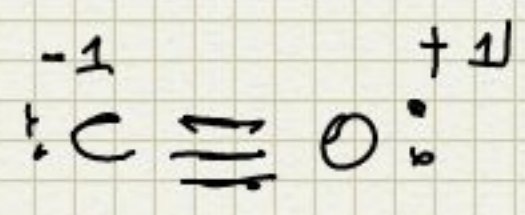


$$6 + 4 = 10$$

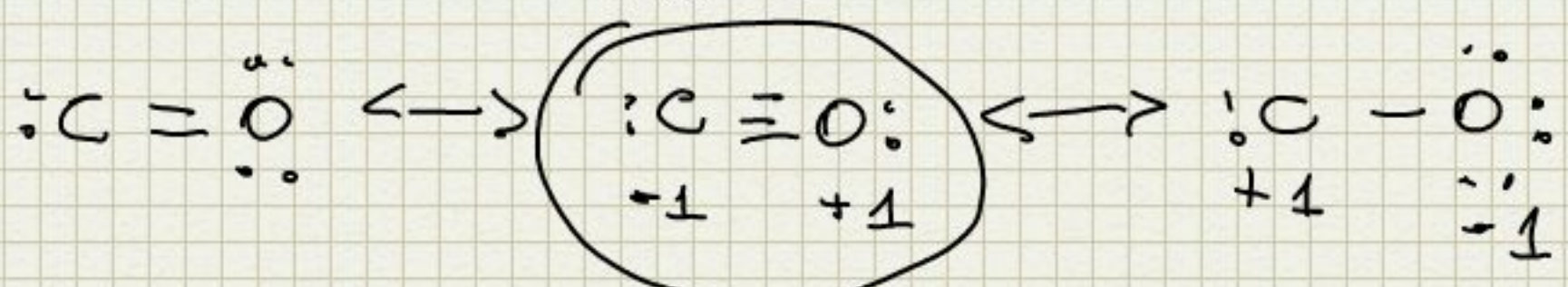
5 COPPIE



OTT.

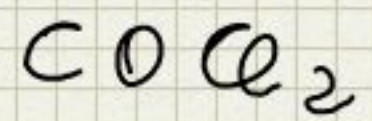


VOLENDO



SPERIMENTALE

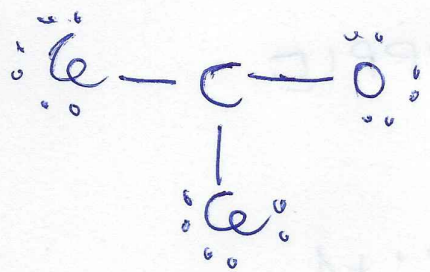
CLORURO DI CARBONILE (FOSGENE)



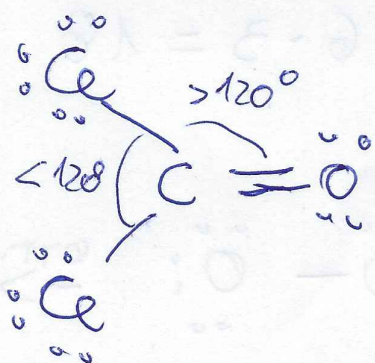


$$7 \cdot 2 + 6 + 4 = 24$$

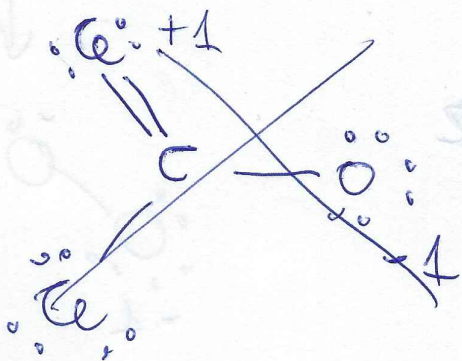
12 COPPIE



OTT.



C IBRIDATO sp^2



NO

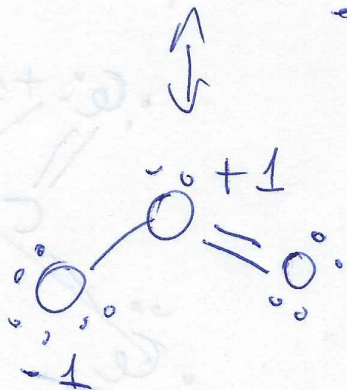
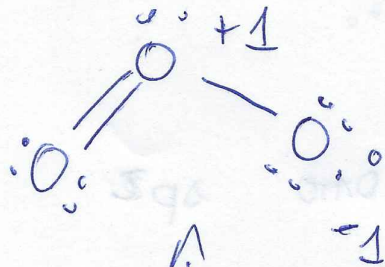
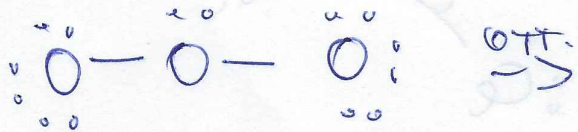
NOX . C = +4

$$\begin{array}{l} 2 - 0 \\ 1 - 2 \end{array}$$

O₂O N O

O₃ M.OX O = 0

$\ddot{O}:$ $6 \cdot 3 = 18$ 9 COPPIE

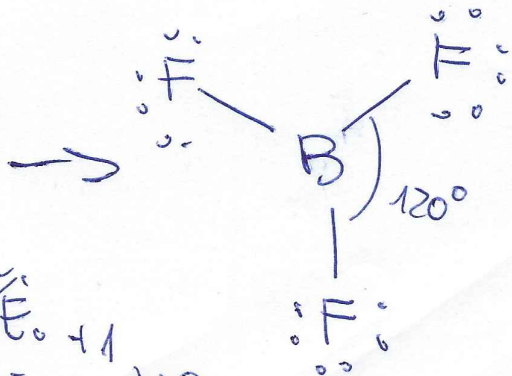
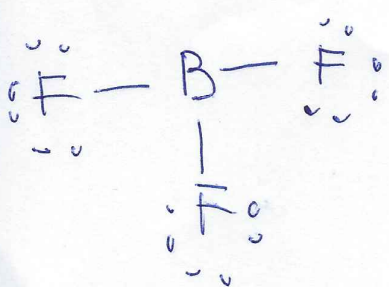


O IBRIDATO sp^2

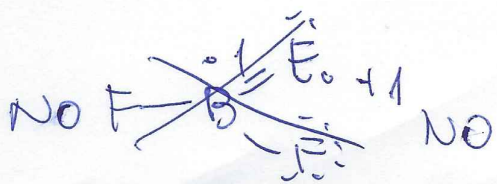
BF₃ TRIFLUORURO DI BORO

M.OX B = +3

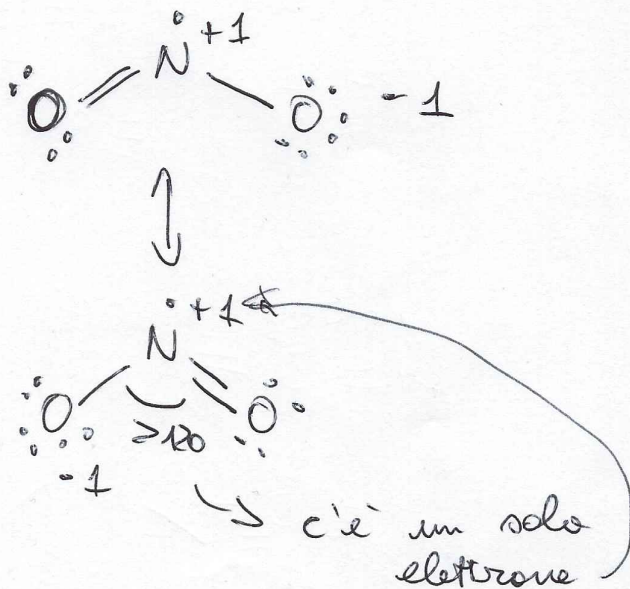
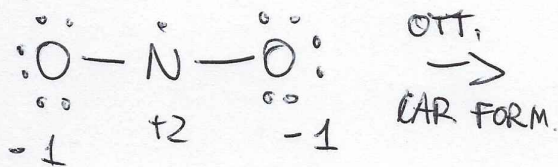
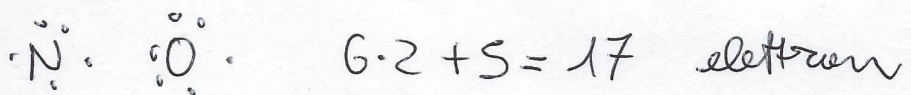
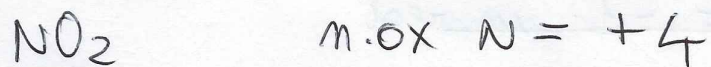
$\ddot{B}:$ $\ddot{F}:$ $7 \cdot 3 + 3 = 24$
12 COPPIE



B IBRIDATO sp^2

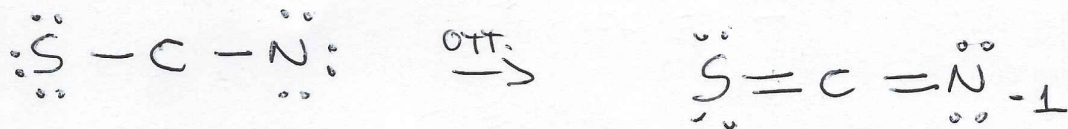
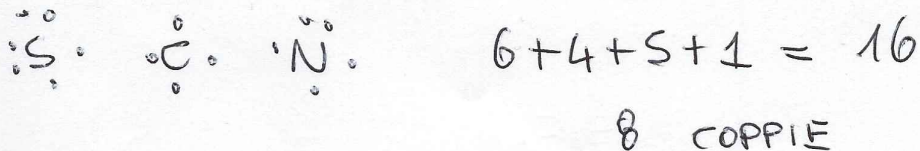
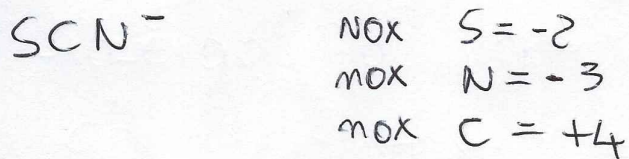


DIOSSIDO DI AZOTO



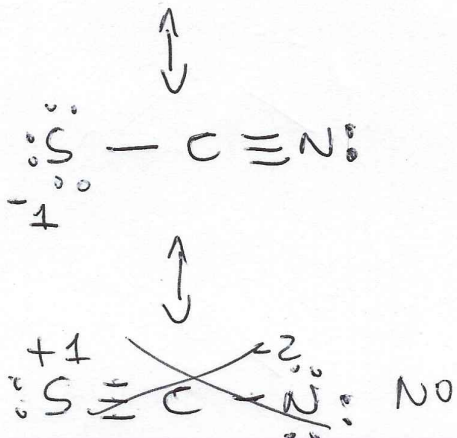
N IBRIDATO sp^2

TIOCIANATO

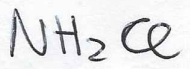


angolo 180°

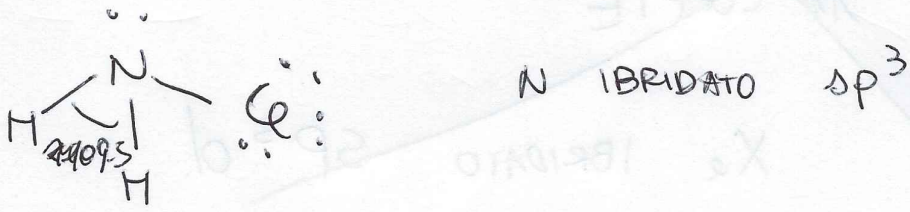
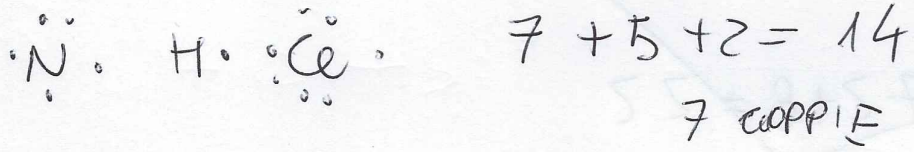
C IBRIDATO sp



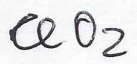
CLOROAMMINA



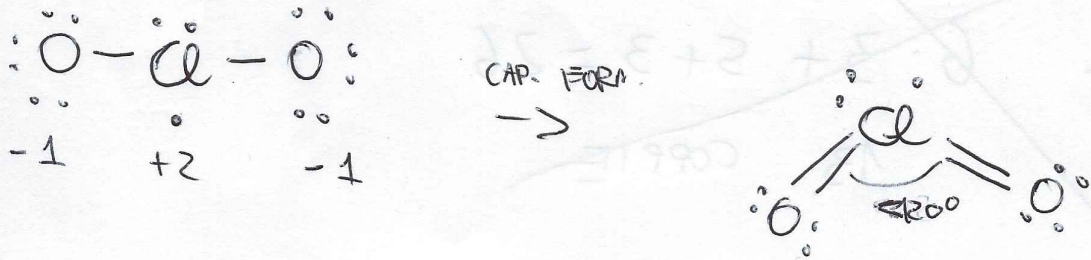
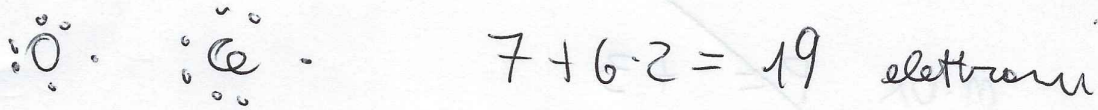
MOX N = -3 Cloroammine.



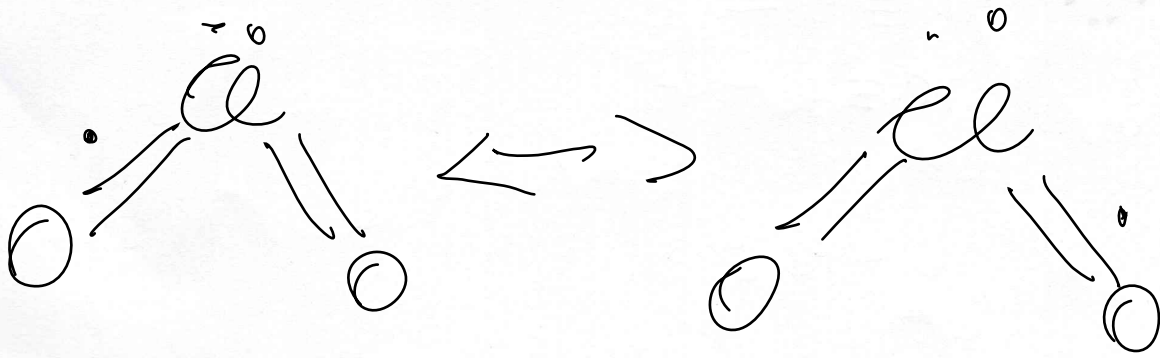
DIOSSIDO DI CLORO



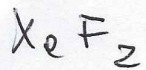
MOX Cl = +4



PROPOSTA DA PAULING.

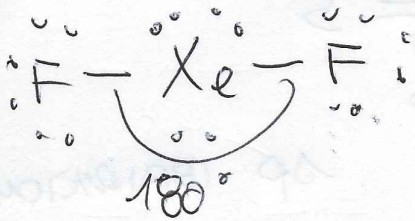


DIFLUORURO DI XENO



$$7 \cdot 2 + 8 = 22$$

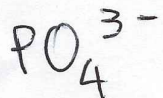
11 COPPIE



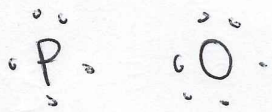
Xe IBRIDATO sp^3d

LINEARE

FOSFATO

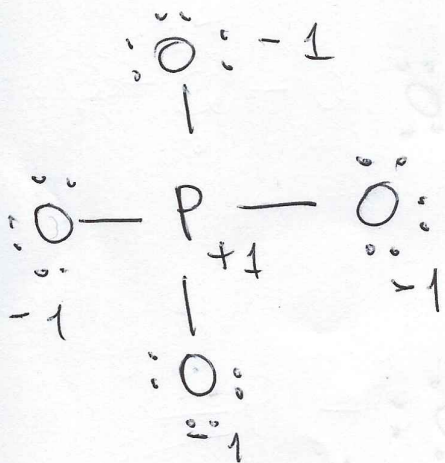


max $\beta = +5$

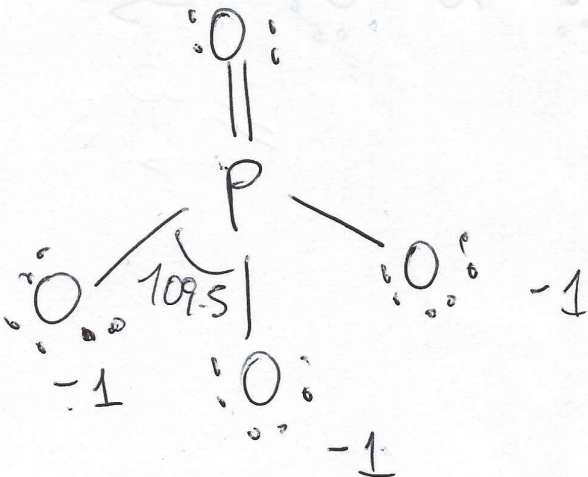


$$6 \cdot 4 + 5 + 3 = 32$$

16 COPPIE

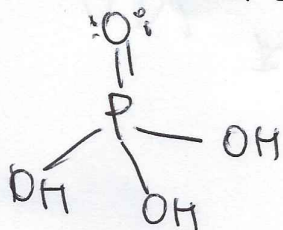


→

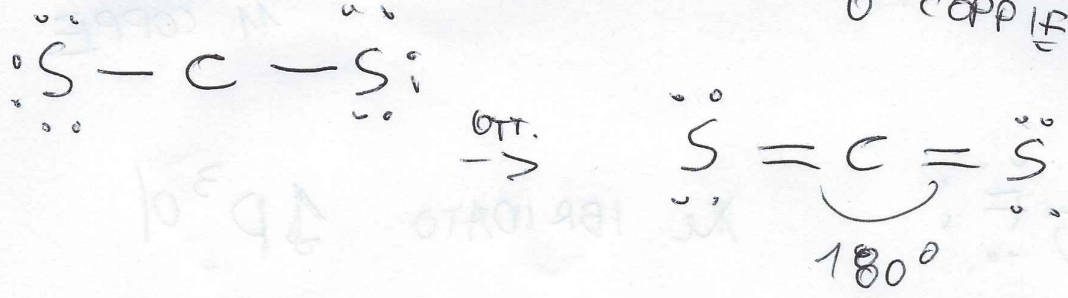
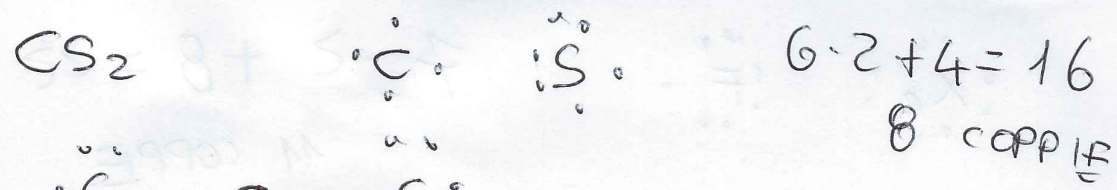


sp^3

ACIDO FOSFORICO

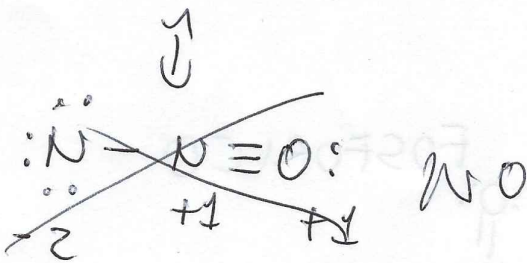
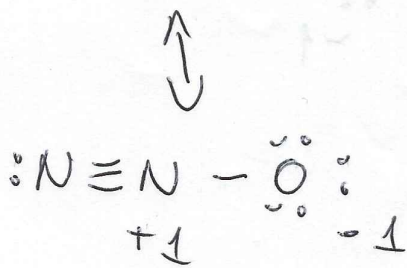
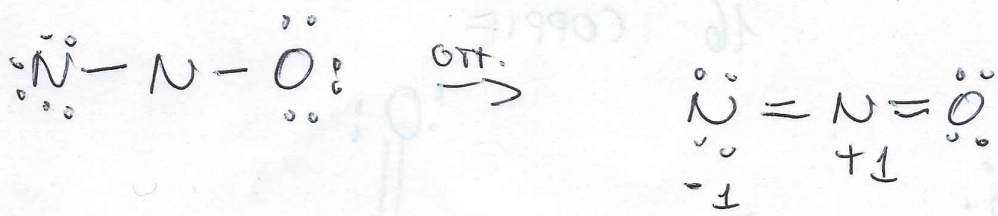


DISOLFURO DI CARBONIO



sp IBRIDAZIONE

MONOSSIDO DI DIAZOTO



NO

CIANATO



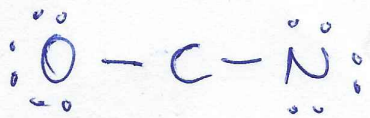
MOX C = +4

MOX N = -3

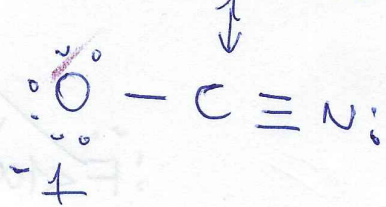
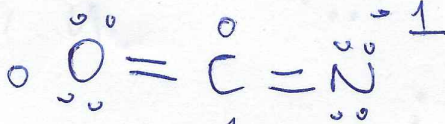


$6 + 4 + 5 + 1 = 16$

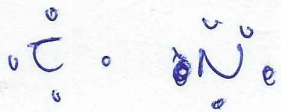
8



OTT. \rightarrow



CIANURO

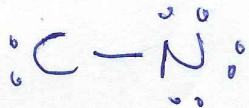


M.OX C = +2

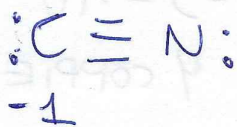
M.OX N = -3

$5 + 4 + 1 = 10$

5 COPPIE



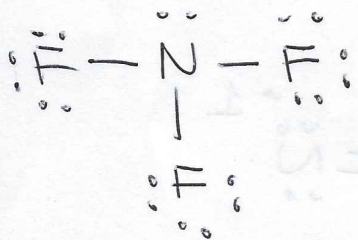
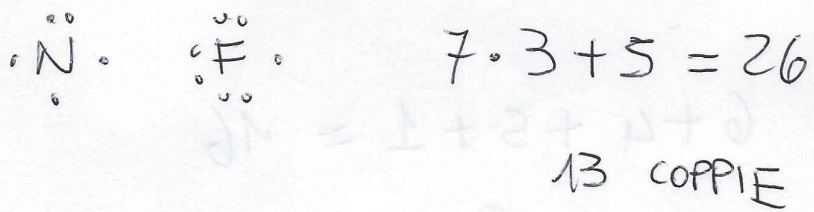
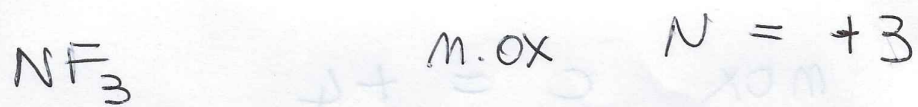
OTT. \rightarrow



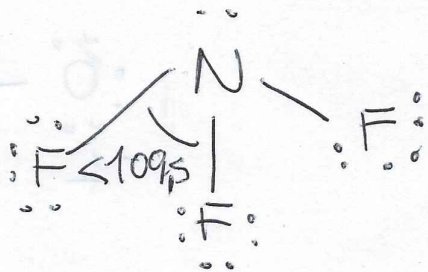
ACIDO CIANIDRICO



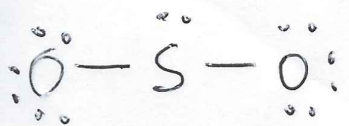
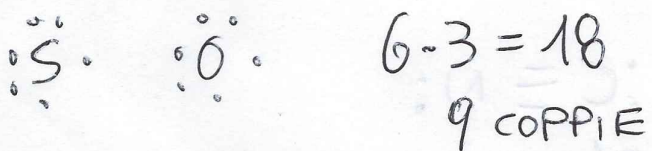
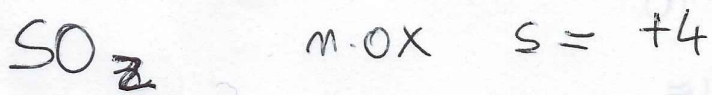
TRIFLUORURO DI AZOTO



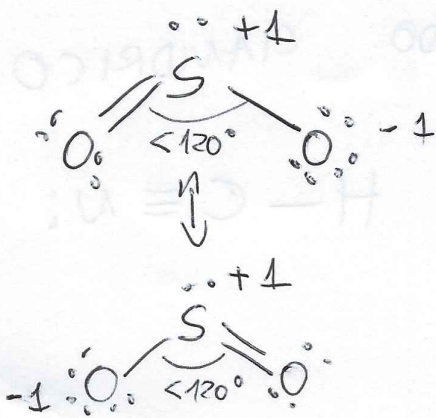
N IBRIDATO sp^3



DISSIDO DI ZOLFO



$6\pi \rightarrow$



S IBRIDATO sp^2